

# V.S.R. Government Degree & P.G. College Movva, Krishna Dt. 521135



NAAC Reaccredited with a CGPA of 3.01 score at 'A' Grade (Affiliated to Krishna University) www.gdcmovva.com Email:gdcjkc.movva@gmail.com

## DEPARTMENT OF CHEMISTRY

#### BRIDGE COURSE FOR FIRST YEAR STUDENTS

#### 16.11.21 TO 25.11.21

The Department of Chemistry conducted a 9 - Day Bridge Course to the newly joined students in various programmes/courses offered by the college, such as B.Sc. MPC,BZC and BTBC during the academic year 2021-2022, with an objective to bridge the gap between the aspects they studied at Pre-University Level and the syllabus they would be studying in Graduation. The syllabus for the course is framed in such a way that the freshers' will get the basic knowledge on the topics they are going to learn in their graduation. The faculty, Department of Chemistry Smt.M.N.Parameswari handled the sessions and helped the students to adapt themselves to the new academic Curriculum.





## ANDHRAPRADESHSTATECOUNCILOFHIGHEREDUCATION

(AStatutorybodyoftheGovernment of AndhraPradesh)

www.apsche.org**Email**:acapsche@gmail.com

# Structure of Chemistry Core Syllabus under CBCS

YEAR	SEMESTER	COURSE	TITLE	MARKS	CREDITS
I	Ι	Ι	Inorganic and PhysicalChemistry	100	03
-			Practical–IAnalysis ofSALTMIXTURE	50	02

#### **BRIDGE COURSE SYLLABUS**

#### I SEM

#### p-BLOCK ELEMENTSGROUP-15 ELEMENTS

Occurance- electronic configuration, atomic and ionic radii, ionisationenergy, electronegativity, physical and chemical properties

#### **dANDfBLOCKELEMENTSBLOCKELEMENTS**

Position in the periodic table

Electronic configuration of the d-blockelements

General properties of the transition elements (d-block) -physicalproperties, variation in atomic and ionic sizes of transition series, ionisationenthalpies, oxidationstates.

#### SOLIDSTATE

Generalcharacteristicsofsolidstate

Amorphousandcrystallinesolids

Classification of crystallines olids based on different binding forces (molecular, ionic, metallic and covalent solids Probing the structure of solids: X-ray crystallography Crystallattices and unit cells

#### SOLUTIONS

Types of solutions

Expressing concentration of solutions- mass percentag, volume percentage, mass by volume percentage, parts per million

mole fraction

molarity and molality

Solubility: Solubility of a solid in a liquid, solubility of agasinaliquid, Henry'slaw

Vapourpressureofliquidsolutions:vapourpressure of liquid- liquid solutions. Raoult's law as a special caseof Henry's law -vapourpressureofsolutionsofsolidsinliquids

Idealandnon-idealsolutions

Colligative properties and determination of molar mass-relative lowering ofvapour pressure- elevation of boiling point-depression of freezing point-osmosisandosmoticpressure-reverseosmosis



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#### FEED BACK

#### **IV SEMESTER V PAPER SYLLABUS**

#### INORGANICCHEMISTRY

#### UNIT-I

#### CoordinationChemistry

IUPACnomenclatureofcoordination compounds,Structural andstereoisomerismin complexes with coordination numbers 4 and 6. Valence Bond Theory (VBT): Inner and outerorbital complexes. Limitations of VBT, Crystal field effect, octahedral symmetry. Crystalfield stabilization energy (CFSE), Crystal field effects for weak and strong fields. Tetrahedralsymmetry, Factors affecting the magnitude of crystal field splitting energy, Spectrochemicalseries, Comparison of CFSE for Octahedral and Tetrahedral complexes, Tetragonal distortionofoctahedralgeometry,Jahn-Tellerdistortion, squareplanarcoordination.

12h

#### UNIT-II

#### 1. InorganicReactionMechanism:

Introduction to inorganic reaction mechanisms.Concept of reaction pathways, transitionstate, intermediate and activated complex. Labile and inertcomplexes, ligand substitution reactions -  $SN^1$  and  $SN^2$ , Substitution reactions in square planar complexes, Trans-effect, theories of transfect and its applications

#### 2. Stabilityofmetalcomplexes:

Thermodynamicstabilityandkineticstability,factorsaffectingthestabilityofmetalcomplexes, chelate effect, determination of composition of complex by Job's method andmoleratio method.

#### **BioinorganicChemistry:**

Metalions present inbiological systems, classification of elements according to their action inbiological system. Geochemical effect on the distribution of metals, Sodium/K-pump, carbonican hydrase and carboxy peptidase.

Excessanddeficiencyofsometracemetals.Toxicityofmetalions(Hg,Pb,CdandAs),reasonsf ortoxicity,Useofchelatingagentsinmedicine,Cisplatinasananti-cancerdrug.Iron and its application in bio-systems, Haemoglobin, Myoglobin.Storage and transferofiron.

#### PHYSICALCHEMISTRY

#### UNIT-III

#### 1.Phaserule

Concept of phase, components, degrees of freedom. Thermodynamic derivation of Gibbs phase rule. Phase diagram of one component system - water system, Study of Phasediagrams of Simple eutectic systems i) Pb-Ag system, desilverisation of lead ii) NaCl-Watersystem, Congruentandincongruentmeltingpoint-

Definitionandexamplesforsystemshavingcongruent and incongruent meltingpoint ,freezingmixtures.

#### **UNIT-IV**

#### Electrochemistry

Specific conductance, equivalentconductance and molar conductance- Definition and effectof dilution.Cell constant. Strong and weak electrolytes, Kohlrausch's law and its applications,Definitionoftransportnumber,determinationoftransportnumberbyHittorf'smethod. Debye-Huckel-

Onsagar's equation for strong electrolytes (elementary treatment only), Application of conductivity measurements-conduct ometric titrations.

4h

2h

# 8h

# 34 h

Electrochemical Cells- Single electrode potential, Types of electrodes with examples: Metalmetal ion, Gas electrode, Inert electrode, Redox electrode, Metal-metal insoluble saltsaltanion.DeterminationofEMFofacell,Nernstequation,ApplicationsofEMFmeasurements -Potentiometrictitrations.

Fuelcells-Basicconcepts, examples and applications

#### UNIT-V

#### ChemicalKinetics:

#### 14h

The concept of reaction rates. Effect of temperature, pressure, catalyst and other factors onreaction rates. Orderandmolecularity of a reaction, Derivation of integrated rate equations for zero, first and second order reactions (both for equal and unequal concentrations ofreactants). Half-life of a reaction.General methods for determination of order of a reaction.ConceptofactivationenergyanditscalculationfromArrheniusequation.TheoriesofReacti Rates: Collision theory and Activated Complex theory of bimolecular on reactions.Comparisonofthetwotheories(qualitativetreatmentonly).Enzymecatalysis-Specificity,

factors affecting enzyme catalysis, Inhibitors and Lock & keymodel. Michaels-Menten equation-derivation, significance of Michaelis-Menten constant.

#### Listof ReferenceBooks

- 1. .Text book of physicalchemistrybyS Glasstone
- 2. Concise Inorganic ChemistrybyJ.D.Lee
- 1. AdvancedInorganic ChemistryVol-IbySatyaprakash, Tuli,Basu and Madan
  - 2. AdvancedphysicalchemistrybyGurudeep Raj
  - 3. Principlesofphysical chemistrybyPrutton andMarron
  - 4. Advancedphysical chemistrybyBahl andTuli
  - 5. InorganicChemistrybyJ.E.Huheey
  - 6. Basic Inorganic ChemistrybyCottonandWilkinson
  - 7. Atextbookofqualitativeinorganicanalysis by A.I. Vogel
- 8. Atkins, P.W.&Paula, J. deAtkin's Physical Chemistry Ed., Oxford University Press 10th Ed(2014).
  - 9. Castellan, G.W. Physical Chemistry 4th Ed. Narosa (2004).
  - 10. Mortimer, R.G. Physical Chemistry 3rd Ed. Elsevier: NOIDA, UP (2009).
  - 11. Barrow, G.M. Physical Chemistry

#### BRIDGE COURSE SYLLABUS IV SEMESTER V PAPER

#### **COORDINATIONCOMPOUNDS**

Werner's theory of coordination compounds Definitions of some termsusedincoordinationcompounds Nomenclatureofcoordinationcompounds-IUPACnomenclature Isomerism in coordination compounds Bonding in coordination compounds. (a)Valence bondtheory (b) Crystal field theory

#### ${\bf ELECTROCHEMISTRYANDCHEMICALKINETICSELECTROCHEMISTRY}$

Electrochemicalcells Galvaniccells:measurementofelectrodepotentials Conductance ofelectrolyticsolutions-measurementoftheconductivityofionicsolutionsvariationofconductivityandmolarconductivitywithconcentration-strongelectrolytesandweakelectrolytes

#### CHEMICALKINETICS

Rateofachemicalreaction Factorsinfluencingrateofareaction orderofareaction,molecularityofareaction

Integrated rate equations-zero order reactions-first order reactions-half life of a reaction, Pseudo first order reaction Temperature dependence of the rate of a reaction-effect of catalyst









#### **I BSc Chemistry Syllabus**

#### **H SEMESTER**

#### **ORGANICCHEMISTRY**

**UNIT-I** 

**RecapitulationofBasicsofOrganicChemistry** 

Carbon-Carbonsigma bonds(AlkanesandCycloalkanes)

12hGeneralmethodsofpreparationofalkanes-

WurtzandWurtzFittigreaction,CoreyHousesynthesis,physical and

chemicalpropertiesofalkanes, Isomerismand its effecton properties, Freeradical substitutions;

Halogenation, conceptofrelative reactivity v/sselectivity. Conformational analysis of alkanes (Conf ormations, relativestability and energy diagrams of Ethane, Propane and butane). General molecular f ormulaeofcycloalkanesandrelativestability, Baeyer strain theory, with energy

Cyclohexane conformations

diagram, Conformationsofmonosubstituted cyclohexane.

#### **UNIT-II**

#### Carbon-CarbonpiBonds(AlkenesandAlkynes)

#### 12

hGeneralmethodsofpreparation, physical and chemical properties. Mechanism E1, E2, E1 cbreacti ons,SaytzeffandHoffmanneliminations,ElectrophilicAdditions,mechanism(Markownikoff/ Antimarkownikoff addition) with suitable examples, synandanti-addition; addition of H<sub>2</sub>, X<sub>2</sub>HX. oxymercuration-

demercuration, hydroboration-oxidation,ozonolysis,hydroxylation, Diels Alderreaction,1,2and 1,4-addition reactions in conjugated dienes.

Reactionsofalkynes; acidity, electrophilicand nucleophilicadditions, hydration to form carbonylco mpounds, Alkylationofterminalalkynes.

#### **UNIT-III**

#### **Benzeneand itsreactivity**

Concept of aromaticity, Huckel's rule - application to Benzenoid (Benzene, Naphthalene) andNon-

Benzenoidcompounds(cyclopropenylcation,cyclopentadienylanionandtropyliumcation) Reactions-Generalmechanismofelectrophilicaromaticsubstitution, mechanismofnitration, Friedel- Craft's alkylation and acylation. Orientation of aromatic substitution-ortho, para and meta directing groups. Ring activating and deactivating groups with examples(Electronic interpretation of various groups like NO<sub>2</sub> and Phenolic). Orientation of (i) Amino, methoxy and methylgroups (ii) Carboxy, nitro, nitrile, carbonyland sulphonic acid groups

(ii) Halogens

(Explanation bytakingminimum of oneexample from each type)

# **GENERALCHEMISTRY UNIT-IV** 1. Surfacechemistryandchemicalbonding Surfacechemistry Colloids-Coagulation of colloids-Hardy-Schulzerule. Stability of colloids, Protection of Colloids, Gold number.

Physicalandchemicaladsorption, Langmuiradsorptionisotherm, applications of adsorption.

2. ChemicalBonding

Adsorption-

12h

24 h

6h

Valence bond theory, hybridization, VB theory as applied toClF<sub>3</sub>,Ni(CO)<sub>4</sub>, Molecular orbitaltheory -LCAO method, construction of M.O. diagrams for homo-nuclear and hetero-nucleardiatomicmolecules (N<sub>2</sub>, O<sub>2</sub>, CO and NO).

#### 3. HSAB

Pearson's concept, HSAB principle & its importance, bonding in Hard-Hard and Soft-Softcombinations.

#### **UNIT-V**

#### Stereochemistryof carboncompounds

Molecularrepresentations-Wedge, Fischer,NewmanandSaw-Horseformulae. Opticalisomerism:Opticalactivity-

wave nature of light, plane polarised light, optical rotation and specific rotation.

Chiral molecules- definition and criteria(Symmetry elements)- Definition of enantiomers and diastereomers – Explanation of optical isomerism with examples- Glyceraldehyde, Lacticacid, Alanine, Tartaricacid, 2,3-dibromopentane.

D,L,R,SandE,Z-configurationwithexamples.

DefinitionofRacemic mixture–Resolution ofracemicmixtures (any3techniques)

#### **List of Reference Books**

Morrison, R. N. & Boyd, R. N. Organic Chemistry, Dorling

Kindersley (India) Pvt. Ltd.(PearsonEducation).

Finar, I. L. Organic Chemistry (Volume 1), Dorling Kindersley (India) Pvt. Ltd. (PearsonEducation).

Finar,

I.L.OrganicChemistry(Volume2:StereochemistryandtheChemistryofNat uralProducts),DorlingKindersley(India)Pvt.Ltd. (PearsonEducation). Eliel, E. L. &Wilen, S. H. Stereochemistry of Organic Compounds; Wiley: London, 1994.Kalsi, P. S. Stereochemistry Conformation and Mechanism; New Age International, 2005.

#### **BRIDGE COURSE SYLLABUS**

ORGANIC CHEMISTRY-SOME BASIC PRINCIPLES AND TECHNIQUES AND HYDROCARBONS

General introduction.Tetravalency of Carbon: shapes of organic compounds.

Structural representations of organic compounds.

Classification of organic compounds.Nomenclature of organic compounds.

Isomerism.Fundamental concepts in organic reaction

mechanisms.Fission of covalent bond.

Nucleophiles and electrophiles.

Electron movements in organic reactions.

Electron displacement effects in covalent bonds.

Types of Organic reactions.

#### CHEMICAL BONDING AND MOLECULAR STRUCTURE

Kossel – Lewis approach to chemical bonding.

Ionic or electrovalent bond - Factors favourable for the formation of ionic compounds-Crystal structure of sodium chloride-General properties of ionic compounds.

Bond Parameters – bond length, bond angle, and bond enthalpy, bond order, resonance-Polarity of bonds dipole moment

Valence Shell Electron Pair Repulsion (VSEPR) theories.Predicting the geometry of simple molecules.

Valence bond theory-Orbital overlap concept-Directional properties of bondsoverlapping of atomic orbitals strength of sigma and pi bonds-Factors favouring the formation of covalent bonds

Hybridisation- different types of hybridization involving s, p and d orbitals- shapes of simple covalent molecules.

Coordinate bond -definition with examples.

Molecular orbital theory – Formation of molecular orbitals, Linear combination of atomic orbitals (LCAO)-conditions for combination of atomic orbitals - Energy level diagrams for molecular orbitals -Bonding in some homo nuclear diatomic









#### **BRIDGE COURSE FOR FIRST YEAR STUDENTS**

#### 14.11.22 TO 29.11.22

The Department of Chemistry conducted a 9 - Day Bridge Course to the newly joined students in various programmes/courses offered by the college, such as B.Sc. BZC and BTBC during the academic year 2022-2023, with an objective to bridge the gap between the aspects they studied at Pre-University Level and the syllabus they would be studying in Graduation. The syllabus for the course is framed in such a way that the freshers' will get the basic knowledge on the topics they are going to learn in their graduation. The faculty, Department of Chemistry Smt.M.NagaParameswari handled the sessions and helped the students to adapt themselves to the new academic Curriculum.





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# StructureofChemistryCoreSyllabusunderCBCS

YEAR	SEMESTER	COURSE	TITLE	MARKS	CREDITS
Ι	Ι	Ι	Inorganic and PhysicalChemistry	100	03
			Practical–IAnalysis ofSALTMIXTURE	50	02

# **BRIDGE COURSE SYLLABUS**

I SEM

p-BLOCK ELEMENTSGROUP-15 ELEMENTS Occurance- electronic configuration, atomic and ionic radii, ionisationenergy, electronegativity, physical and chemical properties

#### **dANDfBLOCKELEMENTSBLOCKELEMENTS**

Position in the periodic table

Electronic configuration of the d-blockelements

General properties of the transition elements (d-block) -physicalproperties, variation in atomic and ionic sizes of transition series, ionisationenthalpies, oxidationstates.

#### SOLIDSTATE

General characteristics of solid state

Amorphousandcrystallinesolids Classificationofcrystallinesolidsbasedondifferentbindingforces(molecular, ionic, metallic and covalent solids Probing the structure ofsolids:X-raycrystallographyCrystallatticesandunitcells

#### SOLUTIONS

Types of solutions

Expressing concentration of solutions- mass percentag, volume percentage, mass by volume percentage, parts per million

mole fraction

molarity and molality

Solubility: Solubility of a solid in a liquid, solubility of agasinaliquid, Henry'slaw

Vapourpressureofliquidsolutions:vapourpressure of liquid- liquid solutions. Raoult's law as a special caseof Henry's law -vapourpressureofsolutionsofsolidsinliquids

Idealandnon-idealsolutions

Colligative properties and determination of molar mass-relative lowering ofvapour pressure- elevation of boiling point-depression of freezing point-osmosisandosmoticpressure-reverseosmosis









#### PRE EXAM CONDUCTED





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#### **BRIDGE COURSE SYLLABUS**

ORGANIC CHEMISTRY-SOME BASIC PRINCIPLES AND TECHNIQUES AND HYDROCARBONS

General introduction.Tetravalency of Carbon: shapes of organic compounds.

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Isomerism.Fundamental concepts in organic reaction

mechanisms.Fission of covalent bond.

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Bond Parameters – bond length, bond angle, and bond enthalpy, bond order, resonance-Polarity of bonds dipole moment

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Valence bond theory-Orbital overlap concept-Directional properties of bondsoverlapping of atomic orbitals strength of sigma and pi bonds-Factors favouring the formation of covalent bonds

Hybridisation- different types of hybridization involving s, p and d orbitals- shapes of simple covalent molecules.

Coordinate bond -definition with examples.

Molecular orbital theory – Formation of molecular orbitals, Linear combination of atomic orbitals (LCAO)-conditions for combination of atomic orbitals - Energy level diagrams for molecular orbitals -Bonding in some homo nuclear diatomic









#### **BRIDGE COURSE FOR FIRST YEAR STUDENTS**

#### 18.08.23 TO 29.08.23

A bridge course for newly admitted students is conducted every year before the commencement of the first semester classes. The main objective of the course is to bridge the gap between subjects studied at Pre-university level and subjects they would be studying in Graduation.Bridge courses help bridge the gap between the curriculums of the previous educational system with the new system, ensuring that students have a developed foundation in the subjects they will encounter in the new system,

thereby minimizing the challenges of knowledge gaps. The Department of Chemistry conducted a 10 - Day Bridge Course to the newly joined students in various programmes/courses offered by the college, such as B.Sc. Chemistry ,B.Sc Botany,B.Sc Computersduring the academic year 2023-2024, with an objective to bridge the gap between the aspects they studied at Pre-University Level and the syllabus they would be studying in Graduation. The syllabus for the course is framed in such a way that the freshers' will get the basic knowledge on the topics they are going to learn in their graduation. The faculty, Department of Chemistry Smt.M.NagaParameswari handled the sessions and helped the students to adapt themselves to the new academic Curriculum.











# V.S.R. GOVERNMENT DEGREE & P.G.COLLEGEMOVVA, KRISHNA DT

# Activity

## **BRIDGE COURSE**

## I SEMESTER BCOM, BA & BSC

# 2021-2022

# Conducted by

Dr. J .Kalpana

Y. Suneetha

Dr. A.V.V.V.Malleswaramma

#### **DEPARTMENT OF ENGLISH**

#### **A Report on Bridge course**

A Bridge course is conducted every year for newly admitted students before the commencement of the first semester classes. They will have an elementary understanding of the advanced subjects that will be taught them in the future. It ensures them of the basic knowledge of the course that they are going to pursue.

**Objectives:** The main objective of the course is to bridge the gap between subjects studied at Pre-university level and subjects they would be studying in Graduation.

**Outcomes:** It gives ample time to the students to study the fundamentals of the course and helps out a smooth shift from the prior course to the existing course. The aim of the bridge course is to promote the lesser adeptness level of the student to a superior aptitude stature.

#### Methodology::

\*Coaching, tutoring and Individual attention.

\*Identifying the slow, moderate and advanced learners by conducting slip tests during the Bridge course.

SYALLABUS:



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<sup>rd</sup>,4<sup>th</sup>and5<sup>th</sup>floors,NeeladriTowers,SriRamNagar,6<sup>th</sup>BattalionRoad,At makur (V),Mangalagiri(M),Guntur-522 503,AndhraPradesh

Web:www.apsche.orgEmail:acapsche@gmail.com

#### **REVISED SYLLABUS OF ENGLISH under (Part – I)UNDER**

#### CBCSFRAMEWORKWITHEFFECTFROM2020-21

#### PROGRAMME:THREE-YEARB.A./B.Sc./B.Com/BCA/BBM/BHM&CT,etc.

Andhra Pradesh State Council of Higher Education, Mangalagiri, GunturDistrict

# **Revised English Syllabus from 2020-21**

# OnwardsUnder ChoiceBasedCreditSystem

# Semester-I EnglishPraxisCourse-I :ACourseinCommunicationandSoftSkills EnglishPraxisCourse-I <u>ACourse in CommunicationandSoft Skills</u>

#### LearningOutcomes

Bytheend of thecourse thelearnerwill be ableto:

- Usegrammar effectivelyinwritingandspeaking.
- Demonstrate the use of good vocabulary
- Demonstratean understatingofwritingskills
- Acquireabilityto useSoft Skills inprofessional and dailylife.
- Confidentlyusethe tools of communication skills

#### I. UNIT:ListeningSkills

- i. ImportanceofListening
- ii. TypesofListening
- iii. Barriersto Listening
- iv. EffectiveListening

#### II. UNIT:SpeakingSkills

- a. SoundsofEnglish: Vowelsand Consonants
- b. WordAccent
- c. Intonation

#### III. UNIT:Grammar

- a) Concord
- b) Modals
- c) Tenses(Present/Past/Future)
- d) Articles
- e) Prepositions

- f) QuestionTags
- g) SentenceTransformation(Voice,ReportedSpeech&DegreesofComparison)

h) ErrorCorrection

#### VSR Govt. Degree & PG College,

Movva, Krishna Dist.

#### Department of English

#### Pre test

Time:1h

#### e:1h Bridge Course 2021-22 Total=10M

#### I. Fill in the blanks with 'a', 'an', or 'the'

a.\_\_\_\_\_ Ganga is holy river.

- b. India is \_\_\_\_\_ SAARC country.
- c. They waited for him for \_\_\_\_\_ hour.
- d. \_\_\_\_\_rich are not always happy.
- e. \_\_\_\_\_ Mr. Shyam wishes to speak to you.
- f. My friend is \_\_\_\_\_ university professor.

#### II Fill in the blanks with suitable prepositions:

6x1/2=3

6x1/2=3

- a. Anupama left \_\_\_\_\_\_ England to meet for friend.
- b. Suguna hung her family photo \_\_\_\_\_\_ the wall.
- c. Who does this bag belong \_\_\_\_\_\_?
- d. The old man distributed hs property his four sons equally\_\_\_\_\_ e).Please, refrain\_\_\_\_\_ drinking alcohol.
- f. Sunil succeeded \_\_\_\_\_\_ starting the car.

#### III. Fill in the blanks with suitable forms of the verbs given in the brackets. 4x1=4

- a. The earth ..... round the sun. (revolve)
- b. It started to rain while we..... tennis. (play)
- c. Renuka ..... to Vizag recently. (go)
- d. Mohan .....a novel by the end of the next year. (write)

## Bridge course abstract 2021- 2022

	S.N	DATE	NAME OF THE	PROGR	SEM	TOPICS	NO. OF
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# **Classes taken during the academic year 2021-2022**

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S.N O	DATE	SEM	FACULTY CONDUCTED	CLASS	TOPIC	NO.OFSTUDENTS ATTENDED
1	12.11.2021	I	Dr.A.V.V.V.Malleswaramma	Bcom,BSc,BA	Pretest	52
2	15.11.2021		Dr.A.V.V.V.Malleswaramma	Bcom,BSc,BA	Verb Tense	42
3	16.11.2021		Dr.j. Kalpana	Bcom,BSc,BA	Prepositions	33
4	17.11.2021		Dr.j.Kalpana,Y.Suneetha	Bcom,BSc,BA	Prepositions,	53512+45
5	18.11.2021		Smt.Y.Suneetha	Bcom,BSc,BA	Pronouns	12
6	18.11.2021		Dr.A.V.V.V.Malleswaramma	Bcom,BSc,BA	Verb tense	47
7	20.11.2021		Dr.J.Kalpana	BSC	Prepositions	12
8	22.11.2021		Dr.A.V.V.V.Malleswaramma	Bcom-Gen	Degrees of	24
9	22-11-		Dr.J.Kalpana	Bcom-CA	Prepositions-	36
10	24.11.2021		Dr.J.Kalpana	B.Com-CA	One word	43
11	24.11.2021		Smt.Y.Suneetha	BSC,BA	Articles	22
12	26.11.2021		Dr.J.Kalpana	Bcom-CA	One word	21
13	26.11.2021		Dr.A.V.V.V.Malleswaramma	BSc (o)	Listening	23
14	27.11.2021		Dr.J.Kalpana	Bcom,BSc,BA	End Test	53
			Smt.Y.Suneetha			
			Dr.A.V.V.V.Malleswaramma			

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(7)	SK. Meharaj Bhanu [Mphi]	SK MEHARAJ EHAN
(8)	P. Roshini [M.E.C.S]	P. Roshini
<u> </u>	M. Naga pavan [Mpc's]	N. Dogopartan
(10)	B. Ugay Bharath [MPW]	B. UppyBlaath.
(1)	D'Sina Noga Raju (MPW)	P. Bin Noge Paju
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(14)	Dischargedon (BLC)	K padmapaani
(15)	P. Romina (Bote)	P. Srihelishmi
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l in the right tense of the verbs given in brackets.	4 ½M
a. When I opened my eyes, I(see/saw) a str	ange sight.
b. Every morning she (get/gets) up early	v and gets ready for work.
c. The headmaster (want/wants) to talk	to you.
d. Jane (live/lives) with her parents.	
e. I (watered/have watered) the plants just n	ow.
f. It (rain) all day.	
g. Where you (go) last summer	?
h. She(go) to Mumbai two years ago.	
i. The (cry) all night.	
. Fill in the blanks with suitable articles:	2 ½ M
a. The world is happy place.	
b. She is accomplished dancer.	
c. Where is hat that I brought fr	om London?
d. Aluminum is useful metal.	
e. He returned after hour.	ALCONTRACTOR OF A CONTRACTOR OF

3. Write a Self Introduction



**3M** 



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